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[Moles To Moles Stoichiometry Worksheet](#)

KEY Review Unit 6

Chemistry 11

6. Given the equation: $C_{12}H_{22}O_{11} + 12O_2 \rightarrow 12CO_2 + 11H_2O + 5638 \text{ kJ}$

a. How much heat is released during the formation of 9.6 moles of CO_2 ?

$$9.6 \text{ mol } CO_2 \times \frac{5638 \text{ kJ}}{12 \text{ mol } CO_2} = \text{Answer } \underline{4510.4 \text{ kJ}}$$

b. How much heat is released during the formation of 0.036 moles of H_2O ?

$$0.036 \text{ mol } H_2O \times \frac{5638 \text{ kJ}}{11 \text{ mol } H_2O} = 18.45 \text{ kJ}$$

Answer $\underline{18.45 \text{ kJ}}$

c. If 1026 grams of $C_{12}H_{22}O_{11}$ are consumed, how much heat is released?

$$1026 \text{ g } C_{12}H_{22}O_{11} \times \frac{1 \text{ mol}}{342.0 \text{ g}} = 3 \text{ mol } C_{12}H_{22}O_{11} \times \frac{5638 \text{ kJ}}{1 \text{ mol } C_{12}H_{22}O_{11}} = \underline{16914 \text{ kJ}}$$

Answer $\underline{16914 \text{ kJ}}$

d. If 23.76 grams of CO_2 are produced, how much heat is released?

$$23.76 \text{ g } CO_2 \times \frac{1 \text{ mol } CO_2}{44.0 \text{ g } CO_2} = 0.54 \text{ mol } CO_2 \times \frac{5638 \text{ kJ}}{12 \text{ mol } CO_2} = \underline{253.71 \text{ kJ}}$$

Answer $\underline{253.71 \text{ kJ}}$

7. Calculate the amount of heat (in Joules) required to warm 350.0 g of water from 30°C to 35°C . (Heat Capacity (C) for H_2O is $4180 \text{ J/kg}\cdot^\circ\text{C}$)

$$\text{Heat} = m \cdot C \cdot \Delta t$$

$$= 0.350 \text{ kg} \times 4180 \frac{\text{J}}{\text{kg}\cdot^\circ\text{C}} \times 5^\circ\text{C} = \underline{7315 \text{ J}}$$

Answer $\underline{7315 \text{ J}}$

8. 35.112 kJ of heat are added to a 500.0 gram sample of water initially at 7°C . Calculate the final temperature of the water sample. Be careful with units!

$$\text{Heat} = m \cdot C \cdot \Delta t$$

$$35,112 \text{ J} = 0.500 \text{ kg} \times 4180 \frac{\text{J}}{\text{kg}\cdot^\circ\text{C}} \times \Delta t^\circ\text{C}$$

$$\Delta t = \frac{35,112}{(0.500 \times 4180)} = \underline{16.8^\circ\text{C}}$$

Answer $t_{\text{final}} = t_{\text{initial}} + \Delta t$
 $t_{\text{final}} = 7 + 16.8 = \underline{23.8^\circ\text{C}}$

Review Unit 6 Page 4 of 4

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CHEM1001 Worksheet 4: Moles and Stoichiometry. Model 1: Balancing Chemical Equations. Chemical equations specify how chemical reactions occur – the ...

Based on the following equation, how many moles of each product are produced when 5.9 moles of Zn(OH)₂ are reacted with H₃PO₄? (You need.

stoichiometry moles to moles worksheet

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Mole Island Diagram. Limiting Reactants. Generic Stoichiometry. Air Bag Design. Water from a Camel. Rocket Fuel. Water in Space. Excess Reactant. Stoichiometry: Mole-Mole Problems. Chemistry II Dr. Slotsky. 1. N₂ + 3H₂ → 2NH₃. How many moles of hydrogen, H₂, are needed to react with 2.0 moles of... Stoichiometry mole to mole worksheet pdf. 1 Fe 1 Atom Fe + 1 Atom S ----> 1 FeS 10 Atom Fe 10 Atom + 10 ----> 10 Atomic FeS 10 Atomic 10 Molecule Fe+ 32.1mg S Its molecular mass is. 60 g/mol. What is its molecular formula? Unit 2: Formula Writing, Naming Compounds, Equations, & Stoichiometry Page 4 (v) How many moles of HCl are used in this reaction? (ICSE 20105 Ans. ... DDD AA5 MOLE CONCEPT AND STOICHIOMETRY WORKSHEET-29 Time 30 Min.

stoichiometry moles to moles worksheet answers

STOICHIOMETRY: MOLE-MOLE PROBLEMS. Name — Ley. 1. N₂ + 3H₂ + 2NH₃. How many moles of hydrogen are needed to completely react with two moles of nitrogen?

stoichiometry 1 moles to moles worksheet

Mole Conversions Worksheet. There are three mole equalities. They are: 1 mol = 6.02 x 10²³ particles. 1 mol = g-formula-mass (periodic table)., Stoichiometry Practice Worksheet. Participants. Find the specific heat (in J/goK) of this metal. See more ideas about science classroom, teaching science, Dey 29, 1399 AP — Chem I Worksheet #38 Stoichiometry (Mole Ratios) ... B From the balanced chemical equation, use a mole ratio to calculate the number of Farvardin 25, 1399 AP — Students are able to convert between moles of products and reactants. ... Handout: Tyler DeWitt Stoichiometry Video Worksheet How many grams of calcium carbonate will be needed to form 4.29 liters of carbon dioxide? n = PV = (1.00 atm)(4.29 L CO₂) = 0.1742 moles CO₂. RT (0.0821 L atm/ ... STOICHIOMETRY MAP FOR CHEMICAL REACTIONS. BALANCED CHEMICAL EQUATION, REACTANTS, PRODUCTS, GIVEN grams, WANTED grams molar mass molar mass. MOLES. MOLES. Reaction Stoichiometry. 1. Na₂SiO₃ (s) + 8 HF (aq) → 2SiF₆ (aq) + 2 NaF (aq) + 3 H₂O (l) a. How many grams of NaF form when 0.500 mol of HF reacts with The purchase of this material entitles the buyer to reproduce worksheets and activities for classroom use only—not ... 62 Stoichiometry: Mole-Mole Problems. 2238193de0

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